

# Chemistry Exam Information Sheet

Constant	Symbol	Value
Avogadro's number	$N_A$	$6.022 \times 10^{23} \text{ mol}^{-1}$
Gas constant	$R$	$0.08206 \frac{\text{L} \cdot \text{atm}}{\text{K} \cdot \text{mol}}$ or $0.08315 \frac{\text{L} \cdot \text{bar}}{\text{K} \cdot \text{mol}}$ or $8.315 \frac{\text{L} \cdot \text{kPa}}{\text{K} \cdot \text{mol}}$ or $8.315 \frac{\text{J}}{\text{K} \cdot \text{mol}}$
Dissociation constant for water	$K_w$	$1.0 \times 10^{-14}$ at $25^\circ\text{C}$
Rydberg constant	$R_H$	$1.097 \times 10^7 \text{ m}^{-1}$
Rydberg constant (in energy units)		$-2.180 \times 10^{-18} \text{ J}$
Mass of electron	$m_e$	$9.109 \times 10^{-31} \text{ kg}$
Charge of electron	$e$	$1.602 \times 10^{-19} \text{ C}$
Faraday's constant	$F$	$9.649 \times 10^4 \text{ C mol}^{-1}$
Speed of light in a vacuum	$c$	$2.998 \times 10^8 \text{ m s}^{-1}$
Planck's constant	$h$	$6.626 \times 10^{-34} \text{ J s}$

## Differential rate laws

$$-\frac{d[A]}{dt} = k[A]^0$$

$$-\frac{d[A]}{dt} = k[A]^1$$

$$-\frac{d[A]}{dt} = k[A]^2$$

## Integrated rate laws

$$[A]_t - [A]_0 = -kt$$

$$\ln[A]_t - \ln[A]_0 = -kt$$

$$\frac{1}{[A]_t} - \frac{1}{[A]_0} = kt$$

## Other Information

$$\text{Quadratic formula: } x = \frac{-b \pm \sqrt{b^2 - 4ac}}{2a}$$

$$\text{Arrhenius equation: } k = Ae^{\frac{-E_a}{RT}}$$

$$0.00^\circ\text{C} = 273.15 \text{ K}$$

THE PERIODIC TABLE of ELEMENTS																			
1 A																	8 A		
1 <b>H</b> 1.008	2 A		8 B										3 A		4 A	5 A	6 A	7 A	2 <b>He</b> 4.003
3 <b>Li</b> 6.941	4 <b>Be</b> 9.012											5 <b>B</b> 10.81	6 <b>C</b> 12.01	7 <b>N</b> 14.01	8 <b>O</b> 16.00	9 <b>F</b> 19.00	10 <b>Ne</b> 20.18		
11 <b>Na</b> 22.990	12 <b>Mg</b> 24.305	3 B	4 B	5 B	6 B	7 B	8	9	10	1 B	2 B	13 <b>Al</b> 26.98	14 <b>Si</b> 28.09	15 <b>P</b> 30.97	16 <b>S</b> 32.07	17 <b>Cl</b> 35.45	18 <b>Ar</b> 39.95		
19 <b>K</b> 39.10	20 <b>Ca</b> 40.08	21 <b>Sc</b> 44.96	22 <b>Ti</b> 47.88	23 <b>V</b> 50.94	24 <b>Cr</b> 52.00	25 <b>Mn</b> 54.94	26 <b>Fe</b> 55.85	27 <b>Co</b> 58.93	28 <b>Ni</b> 58.69	29 <b>Cu</b> 63.55	30 <b>Zn</b> 65.39	31 <b>Ga</b> 69.72	32 <b>Ge</b> 72.59	33 <b>As</b> 74.92	34 <b>Se</b> 78.96	35 <b>Br</b> 79.90	36 <b>Kr</b> 83.80		
37 <b>Rb</b> 85.47	38 <b>Sr</b> 87.62	39 <b>Y</b> 88.91	40 <b>Zr</b> 91.22	41 <b>Nb</b> 92.91	42 <b>Mo</b> 95.94	43 <b>Tc</b> (98)	44 <b>Ru</b> 101.1	45 <b>Rh</b> 102.9	46 <b>Pd</b> 106.4	47 <b>Ag</b> 107.9	48 <b>Cd</b> 112.4	49 <b>In</b> 114.8	50 <b>Sn</b> 118.7	51 <b>Sb</b> 121.8	52 <b>Te</b> 127.6	53 <b>I</b> 126.9	54 <b>Xe</b> 131.30		
55 <b>Cs</b> 132.9	56 <b>Ba</b> 137.3	57 <b>La</b> *	72 <b>Hf</b> 178.5	73 <b>Ta</b> 180.9	74 <b>W</b> 183.8	75 <b>Re</b> 186.2	76 <b>Os</b> 190.2	77 <b>Ir</b> 192.2	78 <b>Pt</b> 195.1	79 <b>Au</b> 197.0	80 <b>Hg</b> 200.6	81 <b>Tl</b> 204.4	82 <b>Pb</b> 207.2	83 <b>Bi</b> 209.0	84 <b>Po</b> (210)	85 <b>At</b> (210)	86 <b>Rn</b> (222)		
87 <b>Fr</b> (223)	88 <b>Ra</b> (226)	89 <b>Ac</b> **	104 <b>Rf</b> (257)	105 <b>Db</b> (260)	106 <b>Sg</b> (263)	107 <b>Bh</b> (262)	108 <b>Hs</b> (265)	109 <b>Mt</b> (266)	110 <b>Ds</b> (281)	111 <b>Rg</b> (280)	112 <b>Cn</b> (285)	113 <b>Uut</b> (284)	114 <b>Fl</b> (289)	115 <b>Uup</b> (288)	116 <b>Lv</b> (293)	117 <b>Uus</b> (294)	118 <b>Uuo</b> (294)		

*	58 <b>Ce</b> 140.1	59 <b>Pr</b> 140.9	60 <b>Nd</b> 144.2	61 <b>Pm</b> (147)	62 <b>Sm</b> 150.4	63 <b>Eu</b> 152.0	64 <b>Gd</b> 157.3	65 <b>Tb</b> 158.9	66 <b>Dy</b> 162.5	67 <b>Ho</b> 164.9	68 <b>Er</b> 167.3	69 <b>Tm</b> 168.9	70 <b>Yb</b> 173.0	71 <b>Lu</b> 175.0
**	90 <b>Th</b> 232.0	91 <b>Pa</b> (231)	92 <b>U</b> 238.0	93 <b>Np</b> (237)	94 <b>Pu</b> (242)	95 <b>Am</b> (243)	96 <b>Cm</b> (247)	97 <b>Bk</b> (247)	98 <b>Cf</b> (249)	99 <b>Es</b> (254)	100 <b>Fm</b> (253)	101 <b>Md</b> (256)	102 <b>No</b> (254)	103 <b>Lr</b> (257)